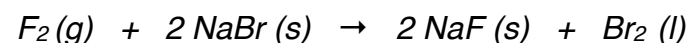
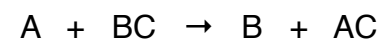


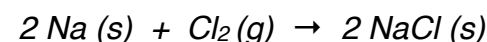
REDOX REACTIONS:

Watch for an element on one side *alone*, but on the other side it is *buddied up with another element*. This implies the element went from a zero charge to either a positive or negative one. There **had** to be a transfer of electrons! These include the following subgroups:

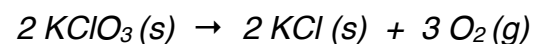
Single displacement *An element and a compound swap one buddy to become another compound and an element.*



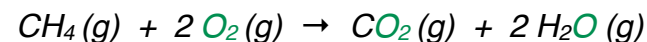
Synthesis (Combination) *Two elements become a compound.*



Decomposition *One substance breaks into its elements.*



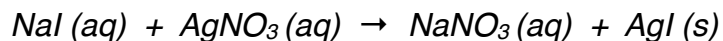
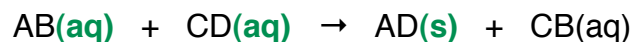
Combustion *Here some substance reacts with oxygen (O₂) to give products with O's stuck on them.*



What to look for:

PRECIPITATION REACTIONS:

Two aqueous solutions get together to form a solid. Often called **double displacement** because two compounds switch buddies.



ACID-BASE REACTIONS:

One reactant has an H⁺, the other an OH⁻. Water is a product. (Another example of double displacement.)

